

## Effect of Synthesized ZnO Nanoparticles As Antioxidant on Cotton Seed Oil and Its Blend with Palm Olein

Research Article

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### Abstract

Nanoparticles of zinc oxide (ZnO NPs) using the precursor citrate was formed as a single phase one stage with the size of 25 nm. ZnO NPs was used at different concentrations (1, 2, 5 mL<sup>-1</sup>) in some edible oils [cotton seed oil and its blend with palm Olein in the ratio of 1:1 (v/v) during heat stress as antioxidant. Some of the chemical characteristics of the oil were estimated as an indicator of deterioration as a result of oxidation during thermal stress within two hours. The results revealed that chemical characteristics (e.g. PV, p-AV and Totox value) of oils that contained 1mL<sup>-1</sup> ZnO NPs were not affected significantly compared to those free from ZnO NPs during heat stress. These results suggested that ZnO NPs has a potential power to control the oxidation process during heat stress of edible oil and could be used as synthetic antioxidant agent. More studies concerning the solubility of ZnO NPs and its nutritional effect on the animal lab are needed.

**Keywords:** Antioxidant; Edible Oil; Heat Stability; Heat Stress; Oxidation; ZnO NPs.

### Introduction

Fried foods spread extensively all over the world as basic meals at fast food outlets. During frying, chemical and physical changes are take place leading to fast deterioration of oils [1]. Most of these changes include oxidation of double bonds in fatty acids [peroxides], hydrolysis of triglycerides producing free fatty acids, pyrolysis and cyclization which effects on the reflective index and color of oil [2]. Deterioration of oil is occurring as a result of either hydrolytic rancidity or oxidative rancidity. Oxidative rancidity refers to reaction of unsaturated fatty acids in lipid with air oxygen to form peroxides which is called induction period [first phase of oil oxidation]. With the continuation of oxidation, peroxides decompose producing mixtures of aldehydes, ketones, acids and other volatile products [secondary phase of oil oxidation]. These products are responsible for "rancid" odors and off-flavors. Also, cis/trans isomerization and polymerization; due to repetition of heating; are increased the melting point and viscosity of the oil [3]. Not only the oil oxidation is altering the nutritive value and flavor of fatty foods, but also toxic substances could be created which considered as health hazard and decreases consumer's acceptability [4]. The oxidation rate can accelerate by high tempera-

ture, presence of water, free fatty acids and catalysts such as heavy metals [5]. Also, oxidation will reduce the storage life and keeping quality of fatty foods. Therefore, antioxidants are used to prevent lipid oxidation and delaying the creation of oxidation products in food system. Free radical scavenging such as Butylated hydroxyanisole [BHA], butylated hydroxytoluene [BHT], propyl gallate [PG], tert-butylhydroquinone [TBHQ] and tocopherols are the most synthetic antioxidants used at the commercial scale to control autoxidation process in fatty foods [6].

With the emergence of nanotechnology, the important properties of nanostructured materials have started motivation among scientists to explore the possibilities of using them in different applications. The search for effective bioactive agents has focused on the development of nanostructures of metals like silver, copper, zinc and gold [6]. ZnO is considered one of the best metal oxides that can be used at a Nano-scale level and currently listed as generally recognized as safe [GRAS] by the U.S. Food and Drug Administration [21CFR182.8991] [FAD 2011]. Also, ZONPs at different concentrations [0.03%, 0.07% and 0.5%] prolong the shelf-life of strawberry fruit in cold storage after harvesting from 0 to 18 days and decreased the microbial load during fruit stor-

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age, delayed weight loss, retained fruit firmness and maintained anthocyanin, vitamin C, phenol content and antioxidant activity at higher concentration than those of untreated fruit [7]. Toxicological studies showed that ZnO NPs could be inhibited lipid peroxidation in cell membrane and internal organs of rats at a concentration of 600mg/ml proving that ZnO NPs may reduce risks to living tissues in both plants and animals [8]. Notice, ZnO NPs has been used for a long time as an active ingredient for dermatological applications in creams, lotions and ointments, depending on its ability to stop the growth of bacteria [9]. Furthermore, it could be used in some foods as antimicrobial. For example, 5 and 8 mM suspensions of ZnO NPs containing citric acid could be effective to stop growth of *Listeria monocytogenes*, *Escherichia coli*, *Staphylococcus aureus* and *Bacillus cereus* in mango juice [10]. ZnO NPs have been contained environmentally safe mineral elements essential to humans when administered in small amount [11].

This work was designed to prepare zinc oxide nanoparticles [ZnO NPs] using the precursor citrate and study its characterization via [the complete name] XRD, [the complete name] TEM and Infra-red spectroscopic [IR] methods. The effect of ZnO NPs at different concentrations on some chemical characteristics of cotton seed oil and its blend with palm Olean in the ratio of 1:1 [v/v] compared to commercial antioxidants [TBHQ] during thermal stress was also studied.

## Material and Methods

Fresh refined, free from synthetic antioxidant, cotton and palm Olean oils as well as TBHQ were kindly supplied from Tanta Oil and Soap Company, Tanta city, Al Gharbia Governorate, Egypt. Zinc nitrate and p-anisidine were purchased from Sigma. All the rest of chemicals used were analytical grade.

### Preparation and characterization of ZnO NPs

Nanocrystalline ZO was prepared by the citrate precursor method according to Duan X, et al., (2005) [12]. Aqueous solutions of Zinc nitrate and citric acid were prepared separate. After that, the mixed solution was dehydrated at 105°C till self-combustion reaction was started within the gel and converted into ash-like powder. The coarse powder was milled, sintered at 650°C with a heating rate of 5°C/min for 2 hours and left to cool at room temperature. The phase identification of the sintered powder was characterized using XRD with X-Ray diffractometer, Philips X'Pert system with Cu-K $\alpha$  radiation [ $\lambda = 0.154056$  nm]. The crystallite size was estimated using Scherrer formula [13] as follows:

$$D = 0.9 \lambda / B \cos \theta$$

Where 0.9 is Scherrer's constant,  $\lambda$  is the wavelength of X-rays,  $\theta$  is the Bragg diffraction angle, and  $\beta$  is the full width at half-maximum [FWHM] of the diffraction peak corresponding to plane |101| located at 36°.

Microstructure and morphology of the powder were characterized by the transmission electron microscope TEM [JEOL, JEM-1400 Electron Microscope] and Fourier transform infrared spectroscopy [FTIR- spectroscopy] [BRUKER Model: Tensor 27] using KBr as sample medium to confirm phase formation.

### Determination of Heat Stability of Oil

About 250 ml of oil samples were placed in 500-ml conical flasks mixed with ZnO NPs at different concentrations [1mM, 2mM and 5mM] along 80 mg TBHQ/kg oil as standard commercial antioxidants were heated with occasionally stirring at 150 rpm and 100°C for different periods [0, 90 180 and 270min]. Fresh refined oil sample free from TBHQ and ZnO NPs, was performed parallel as a control. Samples were withdrawn periodically and subjected to chemical analysis to monitor the oxidation process during heating stress.

### Chemical Analysis

Peroxide value [PV] by iodometric technique[method no 965.33] and p-anisidine value [AnV] [Cd18-90] as described by AOCS [14] of oil samples were determined. The Totox value [TV] was calculated according to Stauffer [15] as follows:

$$\text{Totox value [TV]} = 2PV + p - AV$$

### Rancimat Method

The ability of the oil bearing thermal oxidation was also measured using Schaal oven A 679 Rancimat [Metrohm, Herisan, Switzerland] according to Läubli and Bruttel [16]. About 5g of oil was loaded into the reaction vessel cylinder. The air supply was maintained at 20 ml/min and the heating temperature was kept at 110°C throughout the experiment. Three different samples along control sample were conducted in one batch.

## Results and Discussion

### Characterization of prepared ZO NPs

**XRD:** Figure. 1 represents the X-ray diffraction pattern of ZONPs-powder. A definite line broadening of the XRD peaks indicates that the prepared material consists of particles in nano range. Seven major diffraction peaks were seen in Figure. 1 at 31.4, 34.2, 36.2, 47.4, 56.3, 62.4 and 68.5 regions which can be assigned to diffraction from [100], [002], [101], [102], [110], [103] and [112] planes, respectively. These patterns are compared with those of the standard [JCPDF] cards. Single-phase hexagonal quartzite phase of ZnO NPs structure reveals well-defined diffraction lines and absence of any other phases. The average particle size of the sample was found to be 35 nm which is derived from the FWHM of more intense peak corresponding to 101 plane located at 36.2° using Scherrer's formula.

### Fourier Transform Infrared Spectroscopy [FTIR spectroscopy]

The FTIR spectra of ZnO NPs are helpful in expecting the position of the ions forming the crystal lattice through their vibrational modes. The FTIR spectrum of ZnO NPs was recorded from 200 to 3000 cm<sup>-1</sup>. As shown in Figure. 2, the following bands were observed: a band at 2924 cm<sup>-1</sup> corresponding to stretching vibration of CH<sub>3</sub>; bands at 2924 and 2854 cm<sup>-1</sup> corresponding to antisymmetric and symmetric stretching of C-H [17]; a band at 1505 cm<sup>-1</sup> corresponding to -C H<sub>2</sub> deformation and Zn-O stretching; and a band at 1044 cm<sup>-1</sup> corresponding to C-O stretching peak

Figure 1. Represents the X-ray Diffraction Pattern of ZnO Nanopowder.

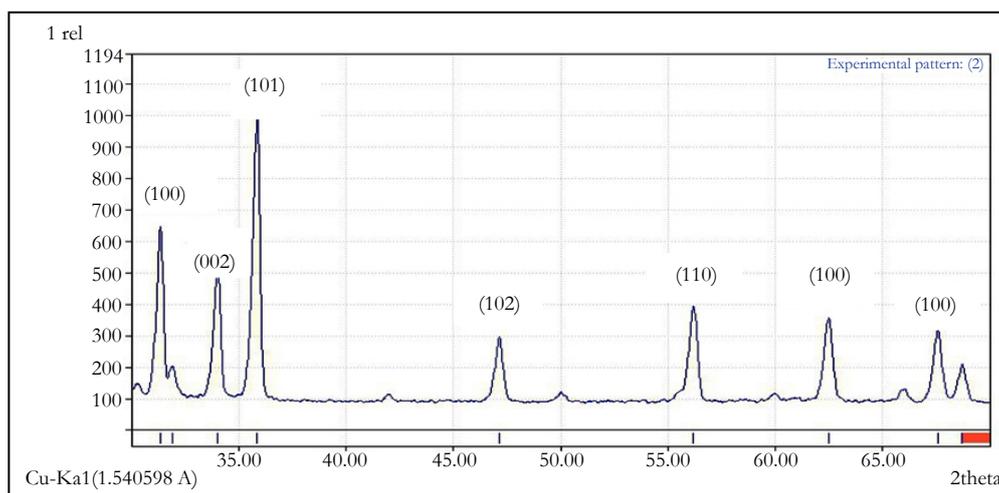
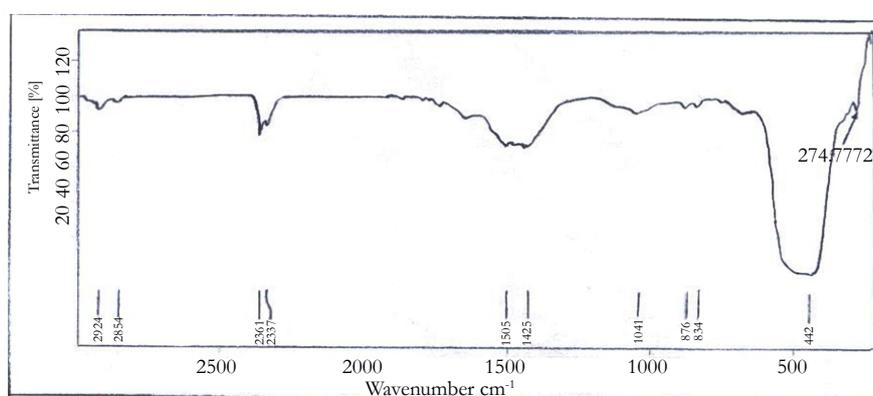


Figure 2. The FTIR Spectrum of ZnO nanoparticles.



[18], and a band at  $834\text{ cm}^{-1}$  corresponding to C-H bending aromatic CH out-of-plane deformation. Parthasarathi and Thilagavathi [19] reported similar FTIR spectra observed of zinc oxide nanoparticles in their investigation.

### TEM Images

It is clear from Figure 3. that TEM images of the ZnO NPs are not exactly circular, and their average size is about 40-50 nm. This value is coinciding to that be calculated from x-ray data, using Scherrer formula.

### Determination of Antioxidant (*Capacity of ZnO NPs in oil*)

**Peroxide value (PV):** The PV is normally used as an index for oil oxidative rancidity. Where, PV increases with the progress of oxidation the double bond in the fatty acids in oil during the induction period. It is clear from Table 1 that PV significantly decreases due to blend cotton oil with palm Olean in the ratio of 1:1 (v/v). The PV of cotton seed oil is 1.50 which decreases to 1.33 for the blend oil. In other words, addition of palm Olean to cotton seed oil decreases PV by 11.33%. This result is in agreement with those reported by Tiwari et al., [20] who stated that mixing oils increases the thermal stability of the oil. Also, the oxidative stability of the oil decreases with extends time of heating. Where the PV of fresh control cotton seed oil is [1.50meqO/kg oil] gradually increases by extended the time of heating stress to reach 12.02 after 270 min. The addition of TBHQ or ZnO NPs to oil increases the

oxidative stability. Where PV of cotton seed oil contains TBHQ or 1mM ZnO NPs after 270 min heating at  $100^{\circ}\text{C}$  are 8.23 and 5.56 meq O/Kg oil, respectively. Also, the results display that the addition of 1mM ZnO NPs seems to be the superior concentration due to the significantly decrease of PV [5.56mequ  $\text{O}_2$ /Kg oil] after 270 min of heating. That is result could be due to their capability of forming excitonic pair [ $e^-$ ,  $h^+$ ] serve as a reactive oxygen species (ROS) scavenger]. The same trend is observed for blend oils, where addition of 1mM ZnO NPs causes a significant decrement in the PV and increases the heat stability of the oil.

**Anisidine Value (AnV):** The AnV is a measure of alpha-beta unsaturated aldehydes or measures high molecular weight saturated and unsaturated carbonyl compounds in tri-acylglycerols of fats and oil [21]. Anisidine value is defined as 100 times of contribution of the extinction in 100 ml solvent and anisidine reagents. As shown in Table 1 heating treatment increases the AnV for most samples after 270 min. The AnV increases in both cotton seed oil and its blend free from antioxidants (TBHQ or ZnO NPs) to reach 44.4 and 43.1 after 270 min heating at  $100^{\circ}\text{C}$ , respectively. The AnV of oils increase to a lesser extent for up to 39.62 and 38.58 in cotton seed oil and its blend containing 1mM ZnO NPs after 270 min heating at  $100^{\circ}\text{C}$  comparing with 35.30 and 34.60 for oil containing TBHQ, respectively. Where AnV of oils contains TBHQ or ZO NPs are significantly lower than that of control sample during the heating stress. It is clear that, as the oil contains TBHQ or ZO NPs; AnV increases in lower rate during the heating stress that of control sample. However, ZnO NPs

shows an excellent result in decreasing AnV. Whereas, adding 1 mL<sup>-1</sup> of ZnO NPs to oil sample provides a decrease in AnV up to 35.30 and 34.60 in cotton seed oil and its blend, respectively. These results are in agreement with those reported by [22] who pointed out that adding of TBHQ to sunflower oil was effective in reducing AnV due to para hydroxyl group which made phenol more easily to donate hydrogen atoms to active free radicals to interrupt the chain reaction of antioxidant. While [23] mentioned that OH<sup>-</sup> in the NPs suspensions was much higher than ones in the bulk particles suspensions.

**Totox Value:** PV and AnV values reflect the oxidation level at early and later stages of oxidation reaction. Therefore, Totox value measures both hydro peroxides and their breakdown products which provide a better estimation of the progressive oxidative deterioration of fats and oils [15]. The data of Table 2 illustrates that Totox value significantly decreases with the addition of 1 mL<sup>-1</sup> ZnO NP's in oil samples. Where, the lowest value is recorded by cotton seed oil (43.00) and its blend (37.20) after 270 min of heating at 100°C. The corresponding values for oil samples containing TBHQ are 51.8 and 48.5, respectively. Although, there is no data on the inhibitory effect of ZnO NP's against lipid oxidation in oil

during heating. This effect was assessed in this study in parallel with convective heating under comparable temperatures. On this topic, there is controversy regarding free radical formation when oils are subjected to heat treatments [24-26] and with the same direction they revealed that even a short period of microwave heating accelerates the formation of some undesirable and harmful compounds (e.g. oxidation products, transformed pigments) due to interactions between electromagnetic field with the chemical constituents of oil.

### Conclusion

In conclusion, ZnO nanoparticles (ZnO NP's) were prepared using citrate precursor method. The XRD, IR analysis, and TEM images clearly indicate the formation of highly pure ZnO nanoparticles. The performance of synthesized ZnO NP's to control the oxidation process of cotton seed oil and its blend during heat stress was tested by different measure parameters and their excellent effect was demonstrated. More studies are need to conduct investigation of the effect of stable suspension of ZnO NP's on vegetable oils and its toxicological experiments.

Figure 3. The TEM Images of ZnO Nanoparticles.

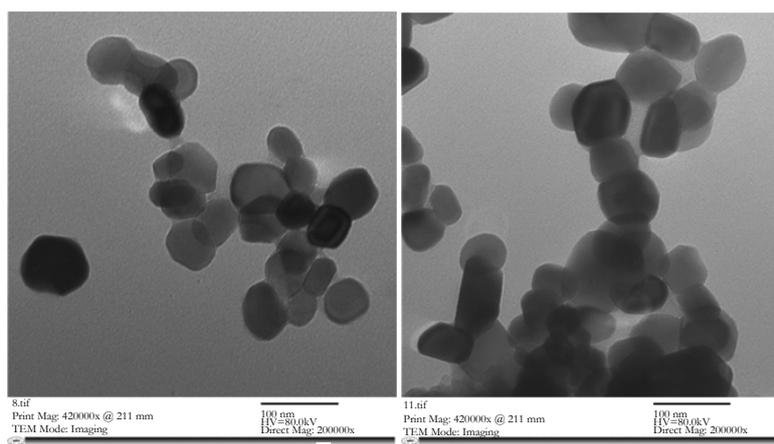


Table 1. *p*-Anisidine value [mg malondialdehyde/kg oil] and Peroxide value [meq O/Kg oil] of oil samples as affected by ZnO nanoparticles at different concentrations as antioxidant during heating stress.

Treatments	Peroxide value [meq O/Kg oil]									
	Cotton seed oil					Blend oil				
	Control	TBHQ	1 mL <sup>-1</sup>	2 mL <sup>-1</sup>	5 mL <sup>-1</sup>	Control	TBHQ	1 mL <sup>-1</sup>	2 mL <sup>-1</sup>	5 mL <sup>-1</sup>
0 min	1.50 ± 0.15	1.50 ± 0.15	1.50 ± 0.15	1.50 ± 0.15	1.50 ± 0.15	1.33 ± 0.02	1.33 ± 0.02	1.33 ± 0.02	1.33 ± 0.02	1.33 ± 0.02
90 min	5.21 ± 0.13	4.46 ± 0.18	3.02 ± 0.21	3.78 ± 0.22	5.43 ± 0.11	5.69 ± 0.13	4.55 ± 0.18	2.73 ± 0.21	3.45 ± 0.77	4.49 ± 0.11
180 min	9.21 ± 0.24	5.67 ± 0.13	4.93 ± 0.19	5.63 ± 0.33	8.57 ± 0.17	7.83 ± 0.24	5.28 ± 0.13	4.17 ± 0.19	5.14 ± 0.33	7.62 ± 0.17
270 min	12.02 ± 0.38	8.23 ± 0.25	5.56 ± 0.32	7.23 ± 0.22	9.95 ± 0.22	9.80 ± 0.38	6.97 ± 0.25	4.81 ± 0.32	6.85 ± 0.22	8.64 ± 0.27
LSD = 0.0611 between treatments					LSD = 0.1473 between treatments					
F = 83.33*** between Time and treatments					F = 76.78*** between time and treatments					
<i>p</i> -Anisidin value [mg malondialdehyde/kg oil]										
0 min	6.33 ± 0.00	6.33 ± 0.00	6.33 ± 0.00	6.33 ± 0.00	6.33 ± 0.00	5.92 ± 0.00	5.92 ± 0.00	5.92 ± 0.00	5.92 ± 0.00	5.92 ± 0.00
90 min	29.10 ± 0.57	24.30 ± 0.59	24.00 ± 0.52	19.50 ± 1.66	18.30 ± 0.31	28.70 ± 0.57	20.90 ± 1.18	19.30 ± 0.52	18.90 ± 1.66	18.00 ± 0.31
180 min	37.50 ± 1.36	35.00 ± 1.28	29.20 ± 0.88	25.10 ± 1.1	23.70 ± 1.29	33.80 ± 1.36	32.60 ± 1.28	25.80 ± 0.88	24.90 ± 1.10	23.10 ± 1.29
270 min	44.40 ± 1.13	39.62 ± 0.74	35.30 ± 0.41	32.20 ± 1.62	31.90 ± 1.21	43.10 ± 1.13	38.58 ± 0.74	34.60 ± 0.46	29.80 ± 1.62	27.60 ± 1.21
LSD = 0.6513 between treatments					LSD = 0.7989 between treatments					
F = 38.13*** between time and treatments					F = 29.75*** between time and treatments					

**Table 2. TOTOX value of oil samples as affected by ZnO nanoparticles at different concentrations as antioxidant during heating stress.**

Treatments	Cotton seed oil					Blend oil				
	Control	TBHQ	1 mL.L <sup>-1</sup>	2 mL.L <sup>-1</sup>	5 mL.L <sup>-1</sup>	Control	TBHQ	1 mL.L <sup>-1</sup>	2 mL.L <sup>-1</sup>	5 mL.L <sup>-1</sup>
0 min	9.30 ± 0.00	9.30 ± 0.00	9.30 ± 0.00	9.30 ± 0.00	9.30 ± 0.00	8.60 ± 0.00	8.60 ± 0.00	8.60 ± 0.00	8.60 ± 0.00	8.60 ± 0.00
90 min	40.90 ± 1.03	32.90 ± 0.88	24.30 ± 0.47	27.10 ± 0.72	35.20 ± 1.33	40.10 ± 1.03	28.40 ± 0.88	23.50 ± 0.47	25.80 ± 0.72	29.90 ± 1.33
180 min	55.90 ± 1.87	40.50 ± 0.67	33.60 ± 0.98	36.40 ± 0.59	52.10 ± 2.40	49.50 ± 1.87	36.40 ± 0.67	31.40 ± 1.01	35.20 ± 0.59	47.80 ± 1.00
270 min	68.40 ± 2.10	51.80 ± 1.04	43.00 ± 1.33	46.70 ± 0.32	59.50 ± 2.04	62.70 ± 2.10	48.50 ± 1.04	37.20 ± 1.33	43.50 ± 0.32	55.90 ± 2.03
LSD = 0.7454 between treatments					LSD = 0.9754 between treatments					
F = 85.90*** between time and treatments					F = 47.27*** between time and treatments					

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